

Tetrakis(3,5-dimethyl-1*H*-pyrazole- κ N²)-(nitrate- κ^2 O,*O'*)cadmium(II) nitrate

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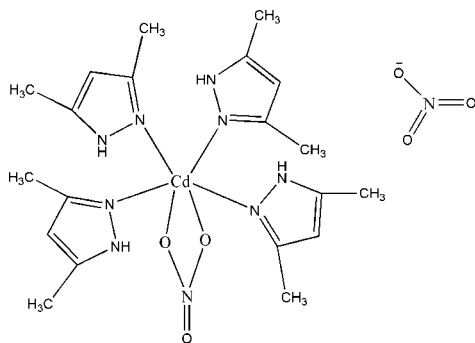
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Key indicators: single-crystal X-ray study; $T = 293$ K; mean $\sigma(\text{C}-\text{C}) = 0.008$ Å; R factor = 0.047; wR factor = 0.137; data-to-parameter ratio = 15.8.

The title compound, $[\text{Cd}(\text{NO}_3)(\text{C}_5\text{H}_8\text{N}_2)_4]\text{NO}_3$, was prepared by reaction of cadmium nitrate and 3,5-dimethylpyrazole in ethanol solution. The Cd atom adopts a distorted *cis*- CdO_2N_4 octahedral geometry involving four dimethylpyrazole molecules and one bidentate nitrate anion. The molecular structure and packing are stabilized by $\text{N}-\text{H}\cdots\text{O}$ and $\text{C}-\text{H}\cdots\text{O}$ inter- and intramolecular hydrogen-bonding interactions.

Related literature

For background on the coordination chemistry of Cd(II) in biological systems, see: Dressing *et al.* (1982). For related literature, see: Addison *et al.* (1984).



Experimental

Crystal data

 $[\text{Cd}(\text{NO}_3)(\text{C}_5\text{H}_8\text{N}_2)_4]\text{NO}_3$
 $M_r = 620.97$

Triclinic, $P\bar{1}$
 $a = 9.1790$ (18) Å
 $b = 11.353$ (2) Å
 $c = 13.669$ (3) Å
 $\alpha = 94.79$ (3)°
 $\beta = 105.61$ (3)°
 $\gamma = 90.68$ (3)°

$V = 1366.2$ (5) Å³
 $Z = 2$
 Mo $K\alpha$ radiation
 $\mu = 0.85$ mm⁻¹
 $T = 293$ (2) K
 $0.25 \times 0.20 \times 0.18$ mm

Data collection

Bruker SMART CCD area-detector diffractometer
 Absorption correction: none
 7482 measured reflections

5035 independent reflections
 4677 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.016$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.046$
 $wR(F^2) = 0.137$
 $S = 1.06$
 5035 reflections
 319 parameters

1 restraint
 H-atom parameters constrained
 $\Delta\rho_{\text{max}} = 1.66$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.87$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{N3}-\text{H4A}\cdots\text{O1}^{\text{i}}$	0.86	2.13	2.943 (8)	157
$\text{N3}-\text{H4A}\cdots\text{O3}^{\text{i}}$	0.86	2.50	3.289 (16)	152
$\text{N4}-\text{H5A}\cdots\text{O1}^{\text{i}}$	0.86	1.92	2.771 (8)	171
$\text{N6}-\text{H7A}\cdots\text{O6}^{\text{ii}}$	0.86	2.33	3.131 (6)	156
$\text{N9}-\text{H10A}\cdots\text{O4}$	0.86	2.50	3.106 (5)	128
$\text{C5}-\text{H5B}\cdots\text{O4}$	0.96	2.59	3.530 (7)	166
$\text{C10}-\text{H10B}\cdots\text{O5}$	0.96	2.29	3.157 (7)	150
$\text{C13}-\text{H13A}\cdots\text{O1}^{\text{iii}}$	0.93	2.49	3.375 (8)	159

Symmetry codes: (i) $x - 1, y, z$; (ii) $-x, -y + 2, -z + 1$; (iii) $-x + 1, -y + 1, -z + 1$.

Data collection: *SMART* (Bruker, 1997); cell refinement: *SAINT* (Bruker, 1997); data reduction: *SAINT*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: AT2668).

References

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 Dressing, S. A., Mass, R. P. & Weiss, C. M. (1982). *Bull. Environ. Contam. Toxicol.* **28**, 172–180.
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supplementary materials

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Tetrakis(3,5-dimethyl-1*H*-pyrazole- κN^2)(nitrate- $\kappa^2 O, O'$)cadmium(II) nitrate

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Comment

It is also known that most of the Cd(II) in biological systems is not in the form of free Cd(II) ions, but is coordinated by the abundance of biological ligands (Dressing *et al.*, 1982). Therefore the coordination chemistry of Cd(II) with ligands is of great interest. In this paper, we reported the synthesis and the crystal structure of tri(3,5-dimethyl pyrazolyl)cadmium(II) nitrate (I).

In the molecule of (I) (Fig. 1), each Cd atoms is coordinated by four nitrogen atoms from four 3,5-dimethyl pyrazoles respectively and two oxygen atoms from nitrate anion. All the bond length and angle are in the normal range. Another nitrate anion exists in the crystal lattice.

The molecular structure and packing are stabilized by the N—H \cdots O and C—H \cdots O inter and intramolecular hydrogen-bonding interactions.

Experimental

Solid 3,5-dimethyl pyrazole 0.96 g (0.01 mol) and cadmium nitrate 0.77 g (0.0025 mol) were added in 50 ml anhydrous alcohol under stirring. The mixture was refluxed for 3.5 h. The colourless solution was filtered and the filtrate was left to stand undisturbed. Upon slow evaporation at room temperature, a colourless crystalline solid appeared three days later and was separated by filtration.

Refinement

H atoms were positioned geometrically and allowed to ride on their parent atoms, with C—H distances of 0.93 and 0.96 Å, and with $U_{\text{iso}}(\text{H}) = 1.2$ or $1.5U_{\text{eq}}$ of the parent atoms.

Figures

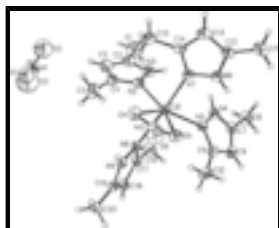


Fig. 1. The molecular structure and atom-labeling scheme for (I), with displacement ellipsoids drawn at the 30% probability level.

Tetrakis(3,5-dimethyl-1H-pyrazole- κ N²)(nitrate- κ ²O,O')cadmium(II) nitrate

Crystal data

$[\text{Cd}(\text{NO}_3)(\text{C}_5\text{H}_8\text{N}_2)_4]\text{N}_1\text{O}_3$	$Z = 2$
$M_r = 620.97$	$F_{000} = 636$
Triclinic, $P\bar{1}$	$D_x = 1.509 \text{ Mg m}^{-3}$
Hall symbol: -P 1	Mo $K\alpha$ radiation
$a = 9.1790 (18) \text{ \AA}$	$\lambda = 0.71073 \text{ \AA}$
$b = 11.353 (2) \text{ \AA}$	Cell parameters from 6067 reflections
$c = 13.669 (3) \text{ \AA}$	$\theta = 2.3\text{--}28.2^\circ$
$\alpha = 94.79 (3)^\circ$	$\mu = 0.85 \text{ mm}^{-1}$
$\beta = 105.61 (3)^\circ$	$T = 293 (2) \text{ K}$
$\gamma = 90.68 (3)^\circ$	Block, colourless
$V = 1366.2 (5) \text{ \AA}^3$	$0.25 \times 0.20 \times 0.18 \text{ mm}$

Data collection

Bruker SMART CCD area-detector diffractometer	4677 reflections with $I > 2\sigma(I)$
Radiation source: fine-focus sealed tube	$R_{\text{int}} = 0.016$
Monochromator: graphite	$\theta_{\text{max}} = 25.5^\circ$
$T = 273(2) \text{ K}$	$\theta_{\text{min}} = 2.3^\circ$
φ and ω scans	$h = -10 \rightarrow 11$
Absorption correction: none	$k = -13 \rightarrow 13$
7482 measured reflections	$l = -16 \rightarrow 11$
5035 independent reflections	

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.047$	H-atom parameters constrained
$wR(F^2) = 0.137$	$w = 1/[\sigma^2(F_o^2) + (0.0885P)^2 + 1.8569P]$
$S = 1.06$	where $P = (F_o^2 + 2F_c^2)/3$
5035 reflections	$(\Delta/\sigma)_{\text{max}} < 0.001$
319 parameters	$\Delta\rho_{\text{max}} = 1.66 \text{ e \AA}^{-3}$
1 restraint	$\Delta\rho_{\text{min}} = -0.87 \text{ e \AA}^{-3}$
Primary atom site location: structure-invariant direct methods	Extinction correction: none

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Cd1	0.13316 (3)	0.80806 (2)	0.33031 (2)	0.03976 (14)
O4	0.3597 (4)	0.8966 (3)	0.4877 (3)	0.0559 (6)
O5	0.1497 (4)	0.9850 (3)	0.4459 (3)	0.0538 (6)
O6	0.3174 (5)	1.0465 (4)	0.5840 (3)	0.0679 (9)
N2	0.2219 (5)	0.6381 (3)	0.2677 (3)	0.0492 (7)
N3	0.1348 (5)	0.5493 (4)	0.2059 (4)	0.0600 (9)
H4A	0.0377	0.5444	0.1918	0.072*
N4	-0.2135 (4)	0.7666 (4)	0.1863 (3)	0.0497 (7)
H5A	-0.2040	0.6919	0.1909	0.060*
N5	-0.1073 (4)	0.8506 (3)	0.2393 (3)	0.0451 (6)
N6	-0.0798 (5)	0.7578 (3)	0.4743 (3)	0.0491 (7)
H7A	-0.1426	0.8045	0.4397	0.059*
N7	0.0527 (5)	0.7270 (3)	0.4544 (3)	0.0495 (6)
N8	0.2600 (4)	0.9168 (3)	0.2410 (3)	0.0461 (6)
N9	0.3644 (4)	1.0038 (3)	0.2868 (3)	0.0448 (7)
H10A	0.4081	1.0120	0.3512	0.054*
N10	0.2779 (4)	0.9758 (3)	0.5082 (3)	0.0464 (7)
C1	0.1461 (12)	0.3654 (6)	0.0977 (7)	0.112 (3)
H1A	0.2229	0.3187	0.0796	0.167*
H1B	0.0800	0.3925	0.0373	0.167*
H1C	0.0887	0.3183	0.1303	0.167*
C2	0.2191 (8)	0.4702 (5)	0.1697 (5)	0.0705 (13)
C3	0.3657 (8)	0.5085 (5)	0.2081 (5)	0.0731 (14)
H3A	0.4503	0.4715	0.1967	0.088*
C4	0.3647 (6)	0.6135 (4)	0.2678 (4)	0.0526 (9)
C5	0.4955 (6)	0.6919 (6)	0.3272 (5)	0.0731 (16)
H5B	0.4594	0.7572	0.3622	0.110*
H5C	0.5480	0.7210	0.2816	0.110*
H5D	0.5632	0.6478	0.3760	0.110*
C6	-0.4658 (7)	0.7409 (7)	0.0621 (6)	0.0848 (19)
H6A	-0.5406	0.7909	0.0249	0.127*
H6B	-0.5090	0.6967	0.1050	0.127*
H6C	-0.4327	0.6872	0.0151	0.127*

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C7	-0.3346 (5)	0.8148 (5)	0.1261 (4)	0.0537 (9)
C8	-0.3053 (6)	0.9339 (5)	0.1395 (4)	0.0543 (9)
H8A	-0.3682	0.9912	0.1079	0.065*
C9	-0.1645 (5)	0.9538 (4)	0.2091 (3)	0.0464 (8)
C10	-0.0833 (7)	1.0697 (5)	0.2516 (5)	0.0689 (14)
H10B	0.0128	1.0559	0.2979	0.103*
H10C	-0.1429	1.1160	0.2873	0.103*
H10D	-0.0676	1.1116	0.1969	0.103*
C11	-0.2363 (8)	0.7287 (5)	0.5941 (5)	0.0694 (13)
H11A	-0.3011	0.7812	0.5520	0.104*
H11B	-0.2042	0.7641	0.6630	0.104*
H11C	-0.2906	0.6553	0.5920	0.104*
C12	-0.1004 (6)	0.7059 (4)	0.5551 (3)	0.0509 (8)
C13	0.0213 (7)	0.6372 (4)	0.5865 (4)	0.0572 (9)
H13A	0.0383	0.5892	0.6401	0.069*
C14	0.1155 (6)	0.6525 (4)	0.5233 (3)	0.0475 (7)
C15	0.2634 (6)	0.5992 (5)	0.5268 (5)	0.0657 (12)
H15A	0.3006	0.6251	0.4726	0.099*
H15B	0.2510	0.5145	0.5189	0.099*
H15C	0.3345	0.6234	0.5912	0.099*
C16	0.1073 (7)	0.8553 (6)	0.0647 (4)	0.0666 (14)
H16A	0.0727	0.7937	0.0979	0.100*
H16B	0.1547	0.8208	0.0150	0.100*
H16C	0.0228	0.8998	0.0314	0.100*
C17	0.2195 (5)	0.9358 (4)	0.1424 (3)	0.0462 (8)
C18	0.2983 (6)	1.0353 (4)	0.1273 (4)	0.0528 (9)
H18A	0.2891	1.0679	0.0658	0.063*
C19	0.3919 (5)	1.0756 (4)	0.2202 (4)	0.0459 (8)
C20	0.5065 (7)	1.1759 (5)	0.2532 (5)	0.0652 (13)
H20A	0.5515	1.1781	0.3254	0.098*
H20B	0.4578	1.2490	0.2373	0.098*
H20C	0.5835	1.1650	0.2180	0.098*
N1	0.7643 (6)	0.4506 (5)	0.1574 (5)	0.0806 (16)
O1	0.8224 (7)	0.5308 (6)	0.2225 (5)	0.1116 (18)*
O2	0.7028 (11)	0.3695 (9)	0.1665 (7)	0.164 (3)*
O3	0.7830 (17)	0.4749 (13)	0.0795 (12)	0.247 (6)*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Cd1	0.0405 (2)	0.0371 (2)	0.0402 (2)	-0.00263 (12)	0.00835 (13)	0.00418 (12)
O4	0.0536 (12)	0.0644 (18)	0.0442 (11)	0.0005 (9)	0.0059 (9)	-0.0028 (10)
O5	0.0618 (15)	0.0429 (11)	0.0496 (14)	0.0006 (11)	0.0051 (9)	-0.0030 (9)
O6	0.079 (2)	0.068 (2)	0.0463 (17)	-0.0047 (17)	0.0045 (13)	-0.0119 (13)
N2	0.0505 (12)	0.0401 (10)	0.0565 (19)	-0.0024 (10)	0.0156 (13)	-0.0010 (10)
N3	0.0675 (15)	0.0397 (16)	0.070 (3)	-0.0119 (14)	0.0175 (19)	-0.0051 (14)
N4	0.0429 (13)	0.0510 (12)	0.051 (2)	-0.0062 (11)	0.0065 (11)	0.0048 (14)
N5	0.0432 (9)	0.0431 (12)	0.0471 (15)	-0.0030 (8)	0.0076 (8)	0.0091 (11)

N6	0.0595 (18)	0.0430 (19)	0.0497 (19)	0.0036 (14)	0.0214 (15)	0.0103 (13)
N7	0.0581 (15)	0.0460 (15)	0.0482 (16)	0.0079 (11)	0.0175 (12)	0.0144 (12)
N8	0.0434 (14)	0.0460 (14)	0.0456 (10)	-0.0075 (11)	0.0050 (10)	0.0097 (11)
N9	0.0458 (18)	0.0451 (18)	0.0430 (12)	-0.0069 (11)	0.0113 (12)	0.0045 (12)
N10	0.0557 (16)	0.0430 (17)	0.0392 (15)	-0.0088 (11)	0.0116 (10)	0.0023 (9)
C1	0.148 (6)	0.059 (4)	0.116 (6)	-0.019 (4)	0.031 (5)	-0.038 (3)
C2	0.093 (2)	0.044 (2)	0.072 (3)	-0.0035 (19)	0.022 (3)	-0.0101 (18)
C3	0.0842 (19)	0.053 (3)	0.088 (4)	0.004 (2)	0.040 (3)	-0.012 (2)
C4	0.0555 (13)	0.045 (2)	0.059 (3)	0.0029 (12)	0.0205 (19)	0.0004 (15)
C5	0.0479 (17)	0.077 (3)	0.086 (4)	0.002 (2)	0.010 (3)	-0.013 (3)
C6	0.054 (3)	0.100 (3)	0.082 (4)	-0.009 (2)	-0.010 (2)	-0.004 (3)
C7	0.0426 (17)	0.0707 (17)	0.045 (2)	0.0005 (13)	0.0080 (12)	0.0024 (18)
C8	0.0503 (19)	0.0663 (15)	0.045 (2)	0.0131 (14)	0.0091 (14)	0.0079 (18)
C9	0.0482 (17)	0.0466 (11)	0.045 (2)	0.0081 (10)	0.0124 (13)	0.0054 (15)
C10	0.071 (3)	0.0418 (14)	0.085 (4)	0.0075 (18)	0.006 (3)	0.005 (2)
C11	0.089 (3)	0.066 (3)	0.064 (3)	-0.009 (2)	0.043 (3)	0.000 (2)
C12	0.074 (2)	0.039 (2)	0.043 (2)	-0.0074 (16)	0.0224 (17)	-0.0003 (14)
C13	0.081 (3)	0.046 (2)	0.044 (2)	-0.0043 (18)	0.0141 (18)	0.0110 (15)
C14	0.0612 (19)	0.033 (2)	0.044 (2)	-0.0021 (14)	0.0061 (15)	0.0065 (13)
C15	0.061 (2)	0.051 (3)	0.078 (4)	0.0071 (19)	0.003 (2)	0.016 (2)
C16	0.067 (3)	0.080 (3)	0.0456 (19)	-0.015 (2)	0.0061 (19)	-0.002 (2)
C17	0.047 (2)	0.053 (2)	0.0399 (11)	-0.0012 (14)	0.0138 (13)	0.0033 (13)
C18	0.060 (3)	0.057 (2)	0.0459 (14)	-0.0009 (16)	0.0198 (16)	0.0114 (15)
C19	0.047 (2)	0.0426 (19)	0.0525 (16)	-0.0011 (13)	0.0213 (15)	0.0070 (14)
C20	0.071 (3)	0.055 (3)	0.076 (3)	-0.0182 (19)	0.032 (2)	0.003 (2)
N1	0.074 (3)	0.075 (3)	0.085 (4)	-0.042 (3)	0.008 (3)	0.012 (3)

Geometric parameters (Å, °)

Cd1—N7	2.278 (4)	C6—C7	1.486 (7)
Cd1—N2	2.293 (4)	C6—H6A	0.9600
Cd1—N5	2.303 (4)	C6—H6B	0.9600
Cd1—N8	2.314 (4)	C6—H6C	0.9600
Cd1—O5	2.427 (3)	C7—C8	1.362 (7)
Cd1—O4	2.673 (3)	C8—C9	1.386 (7)
O4—N10	1.240 (5)	C8—H8A	0.9300
O5—N10	1.264 (5)	C9—C10	1.497 (7)
O6—N10	1.225 (5)	C10—H10B	0.9600
N2—C4	1.343 (6)	C10—H10C	0.9600
N2—N3	1.358 (6)	C10—H10D	0.9600
N3—C2	1.341 (8)	C11—C12	1.500 (7)
N3—H4A	0.8600	C11—H11A	0.9600
N4—C7	1.346 (6)	C11—H11B	0.9600
N4—N5	1.365 (5)	C11—H11C	0.9600
N4—H5A	0.8600	C12—C13	1.363 (8)
N5—C9	1.341 (6)	C13—C14	1.396 (7)
N6—C12	1.350 (6)	C13—H13A	0.9300
N6—N7	1.359 (6)	C14—C15	1.483 (7)
N6—H7A	0.8600	C15—H15A	0.9600

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N7—C14	1.337 (6)	C15—H15B	0.9600
N8—C17	1.335 (6)	C15—H15C	0.9600
N8—N9	1.356 (5)	C16—C17	1.500 (7)
N9—C19	1.342 (6)	C16—H16A	0.9600
N9—H10A	0.8600	C16—H16B	0.9600
C1—C2	1.505 (8)	C16—H16C	0.9600
C1—H1A	0.9600	C17—C18	1.393 (7)
C1—H1B	0.9600	C18—C19	1.366 (7)
C1—H1C	0.9600	C18—H18A	0.9300
C2—C3	1.357 (9)	C19—C20	1.495 (7)
C3—C4	1.389 (7)	C20—H20A	0.9600
C3—H3A	0.9300	C20—H20B	0.9600
C4—C5	1.490 (7)	C20—H20C	0.9600
C5—H5B	0.9600	N1—O2	1.108 (10)
C5—H5C	0.9600	N1—O3	1.176 (15)
C5—H5D	0.9600	N1—O1	1.226 (8)
N7—Cd1—N2	96.81 (14)	C7—C6—H6B	109.5
N7—Cd1—N5	93.80 (14)	H6A—C6—H6B	109.5
N2—Cd1—N5	113.69 (15)	C7—C6—H6C	109.5
N7—Cd1—N8	164.83 (15)	H6A—C6—H6C	109.5
N2—Cd1—N8	89.43 (14)	H6B—C6—H6C	109.5
N5—Cd1—N8	96.33 (13)	N4—C7—C8	105.9 (4)
N7—Cd1—O5	81.49 (13)	N4—C7—C6	121.7 (5)
N2—Cd1—O5	154.81 (14)	C8—C7—C6	132.3 (5)
N5—Cd1—O5	91.50 (13)	C7—C8—C9	107.4 (4)
N8—Cd1—O5	86.98 (13)	C7—C8—H8A	126.3
N7—Cd1—O4	83.20 (13)	C9—C8—H8A	126.3
N2—Cd1—O4	105.38 (13)	N5—C9—C8	109.8 (4)
N5—Cd1—O4	140.88 (12)	N5—C9—C10	122.0 (4)
N8—Cd1—O4	81.82 (12)	C8—C9—C10	128.2 (4)
O5—Cd1—O4	49.43 (12)	C9—C10—H10B	109.5
N10—O4—Cd1	90.4 (2)	C9—C10—H10C	109.5
N10—O5—Cd1	101.7 (3)	H10B—C10—H10C	109.5
C4—N2—N3	105.2 (4)	C9—C10—H10D	109.5
C4—N2—Cd1	128.6 (3)	H10B—C10—H10D	109.5
N3—N2—Cd1	125.3 (3)	H10C—C10—H10D	109.5
C2—N3—N2	111.5 (5)	C12—C11—H11A	109.5
C2—N3—H4A	124.2	C12—C11—H11B	109.5
N2—N3—H4A	124.2	H11A—C11—H11B	109.5
C7—N4—N5	112.0 (4)	C12—C11—H11C	109.5
C7—N4—H5A	124.0	H11A—C11—H11C	109.5
N5—N4—H5A	124.0	H11B—C11—H11C	109.5
C9—N5—N4	104.9 (4)	N6—C12—C13	106.1 (4)
C9—N5—Cd1	130.1 (3)	N6—C12—C11	121.7 (5)
N4—N5—Cd1	123.7 (3)	C13—C12—C11	132.2 (5)
C12—N6—N7	111.5 (4)	C12—C13—C14	107.2 (4)
C12—N6—H7A	124.2	C12—C13—H13A	126.4
N7—N6—H7A	124.2	C14—C13—H13A	126.4
C14—N7—N6	105.9 (4)	N7—C14—C13	109.2 (4)

C14—N7—Cd1	132.7 (3)	N7—C14—C15	122.2 (5)
N6—N7—Cd1	121.3 (3)	C13—C14—C15	128.6 (5)
C17—N8—N9	105.2 (4)	C14—C15—H15A	109.5
C17—N8—Cd1	129.0 (3)	C14—C15—H15B	109.5
N9—N8—Cd1	122.8 (3)	H15A—C15—H15B	109.5
C19—N9—N8	112.3 (4)	C14—C15—H15C	109.5
C19—N9—H10A	123.9	H15A—C15—H15C	109.5
N8—N9—H10A	123.9	H15B—C15—H15C	109.5
O6—N10—O4	123.0 (4)	C17—C16—H16A	109.5
O6—N10—O5	119.2 (4)	C17—C16—H16B	109.5
O4—N10—O5	117.8 (4)	H16A—C16—H16B	109.5
C2—C1—H1A	109.5	C17—C16—H16C	109.5
C2—C1—H1B	109.5	H16A—C16—H16C	109.5
H1A—C1—H1B	109.5	H16B—C16—H16C	109.5
C2—C1—H1C	109.5	N8—C17—C18	109.8 (4)
H1A—C1—H1C	109.5	N8—C17—C16	121.7 (4)
H1B—C1—H1C	109.5	C18—C17—C16	128.5 (4)
N3—C2—C3	106.9 (5)	C19—C18—C17	106.8 (4)
N3—C2—C1	120.8 (6)	C19—C18—H18A	126.6
C3—C2—C1	132.3 (6)	C17—C18—H18A	126.6
C2—C3—C4	106.6 (5)	N9—C19—C18	105.9 (4)
C2—C3—H3A	126.7	N9—C19—C20	121.6 (5)
C4—C3—H3A	126.7	C18—C19—C20	132.5 (5)
N2—C4—C3	109.8 (5)	C19—C20—H20A	109.5
N2—C4—C5	121.6 (4)	C19—C20—H20B	109.5
C3—C4—C5	128.7 (5)	H20A—C20—H20B	109.5
C4—C5—H5B	109.5	C19—C20—H20C	109.5
C4—C5—H5C	109.5	H20A—C20—H20C	109.5
H5B—C5—H5C	109.5	H20B—C20—H20C	109.5
C4—C5—H5D	109.5	O2—N1—O3	124.1 (11)
H5B—C5—H5D	109.5	O2—N1—O1	128.2 (8)
H5C—C5—H5D	109.5	O3—N1—O1	107.6 (9)
C7—C6—H6A	109.5		

Hydrogen-bond geometry (\AA , $^\circ$)

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
N3—H4A \cdots O1 ⁱ	0.86	2.13	2.943 (8)	157
N3—H4A \cdots O3 ⁱ	0.86	2.50	3.289 (16)	152
N4—H5A \cdots O1 ⁱ	0.86	1.92	2.771 (8)	171
N6—H7A \cdots O6 ⁱⁱ	0.86	2.33	3.131 (6)	156
N9—H10A \cdots O4	0.86	2.50	3.106 (5)	128
C5—H5B \cdots O4	0.96	2.59	3.530 (7)	166
C10—H10B \cdots O5	0.96	2.29	3.157 (7)	150
C13—H13A \cdots O1 ⁱⁱⁱ	0.93	2.49	3.375 (8)	159

Symmetry codes: (i) $x-1, y, z$; (ii) $-x, -y+2, -z+1$; (iii) $-x+1, -y+1, -z+1$.

Fig. 1

